give not more than a slight turbidity with Barium Chloride T. S. Allow the magma to drain, then transfer to a suitable vessel and add sufficient water to make 1000 Cc. and mix thoroughly.

In order to obtain a nice white and smooth preparation, one must be careful of the character of the water used. If distilled water is produced in abundance and at a minimum cost, it can be used to advantage. The cost of distilled water to the average pharmacist, however, would preclude its use for the washing of this preparation. Satisfactory water can be cheaply and readily obtained by adding 5 Gm. of powdered Magnesium Carbonate to each litre, boiling and then filtering.

ELIXIR FERRI, QUININAE ET STRYCHNINAE PHOSPHATUM.*

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The formula for the Elixir of the Phosphates of Iron, Quinine and Strychnine, U. S. P. VIII, has been criticised largely because of the uncertainty of the color in different lots and the rapid changes that take place in the color and flavoring on keeping. Recently, another question has been raised, namely, if Quinine in solution with Acetic Acid is not partly changed to Quinotoxin. Consequently, it seems desirable to adopt in the revision a different formula.

The pharmaceutical journals have presented a number of proposed formulas and it has fallen to my lot to try many of these. Without going into a detailed account of the experiments or criticism of these formulas, I will submit the improved formula which I have recommended:

ELIXIR FERRI, QUININAE ET STRYCHNINAE PHOSPHATUM.

Elixir of the Phosphates of Iron, Quinine and Strychnine.

Soluble Ferric Phosphate	17.5	Gm.
Potassium Citrate	5	Gm.
Quinine	8.75	Gm.
Strychnine	0.275	Gm.
Phosphoric Acid	2	Cc.
Alcohol	200	Cc.
Glycerin	200	Cc.
Compound Spirit of Orange	10	Cc.
Purified Talc	30	Gm.
Distilled Water, a sufficient quantity To make	1000	Cc.

Dissolve the Quinine and the Strychnine in the Alcohol and 100 Cc. of Distilled Water to which has been added the Prosphoric Acid. Add to this the Compound Spirit of Orange. Dissolve the Soluble Ferric Phosphate and the Potassium Citrate in 100 Cc. of warm Distilled Water. To this solution add the Glycerin and then the alkaloidal solution and sufficient Distilled Water to make the product measure 1000 Cc. Mix the Purified Talc intimately with the liquid and then filter, returning the first portion of the filtrate until a transparent

* Read before the New Jersey Pharm. Assn., June 11, 1913.

liquid is obtained. Lastly, wash the filter with a mixture of 1 volume of Alcohol and 4 volumes of Water until the filtered product measures 1000 Cc.

In this formula the proportion of the medicinal ingredients is retained the same as in the present official formula, as it was not deemed desirable to make any change in the accepted strength or dosage. The use of glycerin as the sweetening ingredient in place of sugar has proven very satisfactory in elixirs containing iron salts and corrects the tendency of such elixirs to change color. The green tint of the product as at first prepared appears to undergo no marked change after keeping for a year or more. Instead of using Aromatic Elixir as a diluent, the elixir is made in the process of the manipulation, the Compound Spirit of Orange being added, thus insuring the greatest amount of flavoring possible. The manipulation is an important factor in obtaining a satisfactory product and a reversal of the directions as to mixing will promptly demonstrate this.

A RAPID METHOD FOR THE QUANTITATIVE ANALYSIS OF ZINC OINTMENT.*

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In testing some samples of Zinc Oxide Ointment recently, it appeared to me that the analytical methods now in use were too involved and time consuming for the pharmacist: I therefore devised the following very simple and accurate process:

Into a tared porcelain crucible, accurately weigh 1 gram of the sample, heat cautiously until the material bursts into flame, allow to burn quietly until all inflammable material is consumed, then heat strongly with the Bunsen burner until all organic matter is burned off, cool and weigh.

Should difficulty be experienced in burning off organic matter, moisten with a drop of nitric acid, heat cautiously to avoid spattering, and then with the full flame as before.

Since 1 gram of the sample is taken the residue, which is oxide of zinc, can easily be computed into percentage by multiplying the result by 100.

Of course, if necessary this result can be checked by determining the zinc in the residue volumetrically, gravimetrically or electrolytically, and calculating to oxide.

There does not appear to be any reason why the method cannot be employed with equally good results for the analysis of zinc stearate ointment. In this case the amount of stearate present could easily be calculated from the residue of zinc oxide.

The method, in addition to being rapid, is accurate and easily applied.

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* Read before the New York State Pharmaceutical Association, June, 1913.